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Chapter 8 Covalent Bonding Study

1. In covalent bonds, electron sharing

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usually occurs so that atoms attain the electron configuration of noble gases. (8) 2. Atoms form double or triple covalent bonds if they can attain a noble gas structure by sharing two pairs or three pairs of electrons.

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Covalent bond.

covalent. molecule. the same. atoms held together by sharing electrons. in a ... bond, atoms take place in a "tug of war" over electro.... a neutral group of atoms joined together by covalent bonds. all molecules in a sample of a singular molecular compound are....

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molecular compound

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are....
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Chapter 8 Covalent Bonding. chemical bond. covalent bond. molecule. electron dot diagram. □□□the force that holds two atoms together. □□□the chemical bond that results from sharing valence electro.... □□a molecule is formed

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when two or more atoms bond covalently.

□□□□ used to show valence electrons of atoms.

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Chapter 8 Study Guide Covalent Bonding.

Terms in this set (57)

When sharing of electrons occurs, the attachment between atoms that results is

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called a covalent bond. When such an attachment is formed, bond dissociation energy is released, and the process is exothermic.

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Chapter 8/7 Ionic and Covalent Bonding.

covalent bond.
molecule. molecular
substance. diatomic
elements/molecules.

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formed by a shared pair of electrons between atoms. when two or more atoms bond covalently. a substance that is made up of molecules. elements that consist of two atoms covalently bonded.

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Flashcard maker : Lily
Taylor When sharing of
electrons occurs the
attachment between
atoms is called

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...

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Bonding. STUDY. PLAY.
A neutral group of
atoms joined together
by covalent bonds.
molecule. what is a

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molecule called with 2 atoms? diatomic molecule. A tug of war for the electrons takes place between the atoms, bonding the atoms together Atoms held together by SHARING electrons.

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Chapter 8 Covalent Bonding and Molecular Structure 8-11. nuclei. This results in stronger

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attractive forces between electrons and nuclei, decreasing the distance between the nuclei. A carbon-carbon single bond has a bond order of 1 and is longer than a carbon-carbon double bond with a bond order of 2.

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Chapter 8 121 Section

8.1 The Covalent Bond

pages 240–247

Practice Problems page

244 Draw the Lewis

structure for each

molecule. 1. PH₃ H₂ H₂

H— H H P respectively,

for single, double, and

triple P — — 2. H₂ S H

H H — H S S ...

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Covalent Bonding
Single Covalent Bonds
When only one pair of
electrons is shared,
such as in a hydrogen

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molecule, it is a single covalent bond. The shared electron pair is often referred to as the bonding pair.

Chapter 8 Covalent Bonding Study Guide Answers

1. Predict whether the following bonds will be ionic, polar covalent, or non-polar covalent. a. C and O polar- b. Na and Cl ionic c. B and Cl d. H H non 2. Draw the Lewis Dot Diagram for

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the following ionic compounds: a. NaF b. CaCl_2 c. Al_2O_3 3.

Differentiate between an ordinary covalent bond and a coordinate covalent bond. Give an

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Single Covalent Bonds

When only one pair of electrons is shared, such as in a hydrogen

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molecule, it is a single covalent bond. The shared electron pair is often referred to as the bonding pair. For a hydrogen molecule, shown in Figure 8.4, each covalently bonded atom equally attracts the pair of shared electrons.

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A covalent bond in

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which one atom contributes both bonding electrons
Polyatomic ion A tightly bound group of atoms that has a positive or negative charge and behaves as a unit.

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Covalent Bonding •

covalent bond is a shared electron pair that connects the atoms in molecular (covalent) compounds

- when two atoms get close enough, their electron clouds overlap, allowing the nucleus to attract the other atom's electrons which results in a net attraction between the two atoms
- if atoms are too close together,

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they will repulse each

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