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monatomic. Consisting of a single atom.
What is a covalent bond? A type of bond
holding atoms together by sharing
electrons. ... Nonpolar covalent bonds
occur when atoms in a bond pull equally
...

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Change • Chapter 8 121 Section 8.1 The
Covalent Bond pages 240–247 Practice
Problems page 244 Draw the Lewis
structure for each molecule. 1. PH₃ H
H H H— H H P respectively, for single,
double, and triple P — — 2. H₂S H H H
— H S S ...

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a covalent bond between two atoms in
which the shared electron pair comes
from only one of the atoms. ... Chapter 8
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Chapter 8 Covalent Bonding and Molecular Structure 8-4 H₂ molecule. More sophisticated descriptions of chemical bonding will be discussed in Chapter 9. 8.3 Lewis Structures OWL Opening Exploration 8.X One of the most important tools chemists use to predict the properties of a chemical species is its Lewis structure.

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Answer The arrangement is all tetrahedral because all of the central atoms (C, N, S, and F) have the sum of bonding and non-bonding pairs of 4.

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