

## Chapter 9 Mixed Review Stoichiometry Answers

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### Chapter 9 Mixed Review Stoichiometry

CHAPTER 9 REVIEW Stoichiometry SECTION 3 PROBLEMS Write the answer on the line to the left. Show all your work in the space provided. 1. 88% The actual yield of a reaction is 22 g and the theoretical yield is 25 g. Calculate the percentage yield. 2. 6.0 mol of N<sub>2</sub> are mixed with 12.0 mol of H<sub>2</sub> according to the following equation: N<sub>2</sub>(g) + 3H<sub>2</sub>(g) → 2NH<sub>3</sub>(g) N

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CHAPTER 9 REVIEW. Stoichiometry. MIXED REVIEW. SHORT ANSWER Answer the following questions in the space provided. 1. Given the following equation: C<sub>3</sub>H<sub>4</sub>(g) + x. O<sub>2</sub>(g) ( 3CO<sub>2</sub>(g) + 2H<sub>2</sub>O(g) a. What is the value of the coefficient . x. in this equation? b. What is the molar mass of C<sub>3</sub>H<sub>4</sub>? c. How many moles are in an 8.0 g sample of C<sub>3</sub>H<sub>4</sub>? 2. a. What is meant by . ideal conditions

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Stoichiometry b. Theoretically, how many moles of NH<sub>3</sub> will be produced? PROBLEMS Write the answer on the line to the left, Show all your work in the space provided. 1 88% The actual yield of a reaction is 22 g and the theoretical yield is 25 g. Calculate the percentage yield. 2. 6.0 mol of N<sub>2</sub> are mixed with 12.0 mol of H<sub>2</sub> according to the ...

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