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Coordinate Covalent Bond. Covalent Bond. Electronegativity. The region around an atom's nucleus that defines an electron's....  
Forms when one atom donates both of the electrons to be shared.... Chemical bond that results from sharing of electrons.

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STUDY. PLAY. Terms in  
this set (...) chemical  
reaction. the process  
by which the atoms of  
one or more  
substances are  
rearranged to form  
different substances.  
reactants. the starting  
substances in a  
chemical reaction.

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Vocabulary. STUDY.

PLAY. Chemical

Change. A process  
involving one or more  
substances changing  
into a new substance.

Metal. An element that  
is a solid at room  
temperature, a good  
conductor of heat and  
electricity, and is



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generally shiny.

Solution.

Answers

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Change. Terms in this

set (19) chemical

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reaction. the process  
by which the atoms of  
one or more  
substances are  
rearranged to form  
different substances.

**Study 19 Terms |  
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Name CHAPTER

Section 9.1 continued

Date Class STUDY

GUIDE For each of the

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following chemical reactions, write a word equation, a skeleton equation, and a balanced chemical equation. Be sure to show the state of each reactant and product.

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2002. STUDY. PLAY.  
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region around an

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atom's nucleus that  
defines an electron's  
probable location.  
Coordinate Covalent  
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Answers

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**9 Assessment**

**Answers**

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covalent bond.

endothermic reaction.

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exothermic reaction.

Lewis structure.

chemical bond that  
results from the

sharing of valence

electro.... chemical

reaction or process in

which a greater

amount of ener.... a

chemical reaction or

process in which more

energy is release....

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Key**

The products of this reaction are aqueous calcium chloride, carbon dioxide, and water. (s)

+ + 142-0 (Q) 21

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STUDY GUIDE Section

9.2 Classifying

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combustion,  
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replacement reactions.

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Solutions Manual  
Chemistry: Matter and  
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145 CHAPTER 9  
SOLUTIONS MANUAL

35. Aqueous solutions  
of potassium iodide  
and silver nitrate are  
mixed, forming the  
precipitate silver  
iodide. chemical

equation:  $KI(aq) + AgNO_3(aq) \rightarrow AgI(s) + KNO_3(aq)$

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$3(aq) + 0 KNO_3(aq)$

$AgI(s)$  complete ionic

equation:  $K(aq) + I(aq)$

$32 Ag(aq) + NO_3(aq) + 0$

$K(aq) + NO_3(aq) + AgI(s)$

**Chemical**

**Reactions**

**Reactions**

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