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## Haidanore **Structures**

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Structures tr I Name Date: Hour:

Introduction Questions Draw the Lewis structures for the following molecules.

Part c is done for you. a) H<sub>2</sub>O (oxygen is central) b) SO<sub>2</sub> c) N<sub>2</sub> It t-t (a [Jrt @ N=N

@ @ f\ @ 2 In question 1 above you should see three different kinds of bonds

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of bonds as best you can. a) single bond:  
one electron (e-) shared b) double  
bond:

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Description.

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This video is designed to help you complete the Lewis Structures Chemquest. Remember to write down the definitions and take down the notes. Circle and write down any questions you have so you do ...

## **Unit 7: Lewis Structures Chemquest**

24 ChemQuest 8 Name: \_\_\_\_\_ Date:

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Hour: \_\_\_\_\_ Information: Structure of the Atom Note the following symbols: (they are not to scale) = proton (positive charge) = electron (negative charge) = neutron (no charge) The following three diagrams are hydrogen atoms: 1 1H 2 1H 3 1H The following three diagrams are carbon atoms: Notice the type of notation used for atoms: 12 13 14 6C 6



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growing in popularity.

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included in the Lewis Structure.  $4 + 3(6) + 2 = 24$  (carbon has four and each oxygen has six; add two for the -2 charge)  $5 + 3(1) = 8$  (nitrogen has five; each hydrogen has one) Step #2: Draw

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the “skeleton structure” with the central atoms and the other atoms, each connected with a single bond. Step #3: Add six more electron dots to each atom except the

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Pathways (first half videos on ionic  
compounds) Video Tutorials (Tyler  
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## **Dot Structures Answers**

Draw Lewis structures for covalent compounds. The following procedure can be used to construct Lewis electron structures for more complex molecules and ions: How-to: Constructing Lewis electron structures. 1. Determine the total number of valence electrons in the molecule or ion. Add together the

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valence electrons from each atom.

## **10.5: Writing Lewis Structures for Covalent Compounds ...**

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