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Temperature (C) 20 50 100 150 200 250 k (mol⁻¹ dm³ s⁻¹)
0.0147 0.0675 0.497 2.29 7.61 20.1 1/T (K-1) ln k Gradient =
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1 Acid = H₂O, base = NH₃
2 Acid = HCl, base = H₂O
3 Acid = HCOOH, base = KOH
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Kp A) © wwwCHEMSHEETScouk 3-Mar-2018 Chemsheets A2
1156 Nitrogen reacts with hydrogen as shown: $3\text{H}_2(\text{g}) + \text{N}_2(\text{g}) \rightleftharpoons 2\text{NH}_3(\text{g})$ $\Delta H = -76 \text{ kJ mol}^{-1}$ 100 moles of

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Formula Sketch Shape Bond angles Metal oxidation state Co-ordination number *CHO I 1) Draw the two geometric isomers of square planar complex

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TASK 1 - Drawing complexes

reacting mass calculations 1 1) Aluminium is extracted from aluminium oxide as shown. Calculate the mass of aluminium that can be formed from 1020 g of aluminium oxide.

REACTING MASS CALCULATIONS 1 - chemsheets.co.uk | 1pdf.net

initial moles 1 1 1 change in moles -0.8 +0.8 +0.8 equilibrium moles 0.2 1.8 1.8 7) $\text{CH}_4(\text{g}) + \text{H}_2\text{O}(\text{g}) \rightleftharpoons \text{CO}(\text{g}) + 3\text{H}_2(\text{g})$ initial moles 1 1 1 1 change in moles -0.2 -0.2 +0.2 +0.6

EQUILIBRIUM QUANTITIES 1 - WordPress.com

1) Increase the temperature Number of particles with E Ea more same less 2) Remove half the molecules Number of particles with E Ea more same less Energy 4) Reduce the volume of the container Number of particles with E Ea 3) Add a catalyst Number of Energy particles with E Ea more same less Energy 6)

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Add an inert gas more am less Energy

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4.1 Kinetics - A-Level Chemistry

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TASK 2 1) The IR spectra of six compounds are shown. The compounds are: butanoic acid butanone but-3-en-1-ol

Chemsheets A2 029 (Spectroscopy) - Weebly

1) Propene reacts with HBr to form H. H reacts with sodium hydroxide to form I, and I reacts with warm acidified potassium dichromate (VI) to form J. The infra-red spectra of H, I and J are given below, but it does indicate which is - which.

IR TASK 1 - Weebly

Full worked solutions are available to subscribers of www.chemsheets.co.uk. Subscribe for many more exercises with answers. TASK 1 - Bronsted-Lowry acids & bases 1 Acid = H₂O, base = NH₃ 2 Acid = HCl, base = H₂O 3 Acid = HCOOH, base = KOH

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Example 1 Imagine a reaction where R reacts with S to make some products: $R + S \rightarrow \text{products}$ If we compare experiments 1 and 2:

Experiment	Initial [R] mol dm ⁻³	Initial [S] mol dm ⁻³
1	1.0	2.0
2	2.0	2.0

The concentration of R has been doubled, but the concentration of S remains the same. The rate has doubled.

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3) In an experiment, 1.56 g of propan-1-ol ($\text{CH}_3\text{CH}_2\text{CH}_2\text{OH}$) was completely burned in air. The heat evolved raised the temperature of 0.250 dm³ of water from 292.1 K to 339.4 K. Use this data to calculate the enthalpy of combustion of propan-1-ol (the specific heat capacity of water is 4.18 J g⁻¹ K⁻¹).

Chemsheets AS 029 (Thermodynamics) - WordPress.com

This worksheet is designed to accompany chapter 14 in Chemistry: A Molecular Approach. It is a combination of other kinetics worksheets on LibreTexts.

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A rate is a measure of how some property varies with time. Speed is a familiar rate that expresses the distance traveled by an object in a given amount of time. Wage is a rate that

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represents the amount of money earned by a person working for a given amount of time.

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