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Chapter 10 Study Guide The Mole Answer Key.pdf chapter 10 study guide: the mole chapter 10 study guide: the mole matching match each item with the correct statement below. a. molar volume b. molar mass c. atomic mass

___ 1. the number of grams of an element that is numerically equal to the atomic mass of the element in amu ___

2. the mass of

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Stoichiometry and the Mole Study Guide Chemists must have an understanding and a unit of measurement to answer "How much is there?" The mole (6.022×10^{23}) is a unit of measurement that describes matter just as a gross or case describes a quantity of consumer products.

Stoichiometry and the Mole Study Guide - Ms. Osawaru

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The mole is the most common unit used to express the quantity of a chemical substance. For all solids, liquids, and gases, you can convert mass to moles (or moles to mass). For gases, you should memorize the following conversion of volume to moles (or moles to volume): at 0°C and 1 atmosphere

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pressure (known as standard temperature and pressure or STP), one mole of any gas occupies ...

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and contains 6.022×10^{23} atoms.

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